

Mar 7-3:51 PM

Metals usually have 1, 2, or 3 valence (and shell) electrons and tend to lose them to form positive ions, called <u>cations</u> .
Non-metals usually have
5, 6, or 7 valence electrons and tend to <u>gain</u> to form <u>pegative</u> ions, called <u>anions</u> .

- \* Rows are called <u>Periods</u>.
- \* Columns are called <u>Groups</u> or <u>Families</u>.
- \* Elements in the same group share similar, but not exact, characteristics.

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**GROUP 1 - ALKALI METALS** 

- \* Does not include hydrogen!
- \* Shiny, silvery metals.
- \* Form compounds that are white solids and are soluble in water. Ex: NaCl (salt)
- \* Have 1 valence electron
- \* Form 1+ cations



## **GROUP 2 - ALKALINE EARTH METALS**

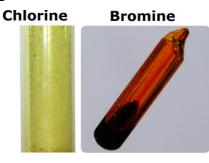


- \* Also shiny, silvery metals.
- \* Form compounds that are often insoluble in water.
- \* Have 2 valence electrons
- \* Form 2+ cations

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## **GROUP 17 - HALOGENS**

- \* All are poisonous
- \* React readily with alkali metals. Ex: KBr, NaCl
- \* Have 7 valence electrons
- \* Form 1- anions





## **GROUP 18 - NOBLE GASES**

- \* Generally do not form compounds.
- \* They are already stable. (have 8 valence electrons)



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Groups 1, 2, and 13-18 are known as the representative elements.

Groups 3-12 in the centre of the periodic table are known as the transition metals.

Metalloids are found between the metals and non-metals because they share properties with both metals and non-metals.

Element Name	Symbol	Group/Family Number	Period Number	Metal, non-metal, or metalloid	Group/Family Name
Chlorine	CI	17	3	Non-Metal	Halogen
Magnesium	Mg	2	3	Metal	Alkaline Farth M
Helium	He	18	1	Non-metal	Noble Gas
Nitrogen	N	12	N	Non-metal	Representative E.
Todine	H	17	5	Non-metal	Halogen
Sodium	Na	1	3	Metal	Alkali Metal
Mercury	Ha	12	6	Metal	Transition M.
Silver	Ag	11	5	Metal	Transition M.
Silicon	S:	14	3	Metalloid	Representative E.
Potassium	K	1	4	Metal	Alhal: Metal

Mar 31-11:08 AM

Noble Gas.jfif

Chem joke.jfif